Metal vs. ligand protonation and the alleged proton-shuttling role of the azadithiolate ligand in catalytic H₂ formation with FeFe hydrogenase model complexes†

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Electron and proton transfer reactions of diiron complexes [Fe₂adt(CO)₆] (1) and [Fe₂adt(CO)₆(PMe₃)₂] (4), with the biomimetic azadithiolate (adt) bridging ligand, have been investigated by real-time IR- and UV-vis-spectroscopic observation to elucidate the role of the adt-N as a potential proton shuttle in catalytic H₂ formation. Protonation of the one-electron reduced complex, 1⁻, occurs on the adt-N yielding 1H and the same species is obtained by one-electron reduction of 1H⁺. The preference for ligand vs. metal protonation in the Fe₂(adt)(I,0) state is presumably kinetic but no evidence for tautomerization of 1H to the hydride 1H₂ in the second coordination sphere. The amine protonation occurs on the adt-N yielding a hydride and a proton. This shows that the adt ligand does not work as a proton relay in the formation of hydride intermediates in the reduced catalyst. A hydride intermediate 1H₂H⁺ is formed only by protonation of 1H with stronger acid. Adt protonation results in reduction of the catalyst at much less negative potential, but subsequent protonation of the metal centers is not slowed down, as would be expected according to the decrease in basicity. Thus, the adtH⁺ complex retains a high turnover frequency at the lowered overpotential. Instead of proton shuttling, we propose that this gain in catalytic performance compared to the propyldithiolate analogue might be rationalized in terms of lower reorganization energy for hydride formation with bulk acid upon adt protonation.

Introduction

Proton binding sites in the second coordination sphere of catalytic metal centres are widely considered as an important design principle for facilitating proton coupled electron transfer in molecular catalysts for e.g. water splitting,¹⁻³ H₂ formation and activation⁶⁻¹² or CO₂ reduction.¹³⁻²¹ For the directed design of synthetic catalysts it is, however, important to note that improvements to catalytic performance (TOF, overpotential) brought about by basic sites in the second coordination sphere are not necessarily arising from a proton relay activity. Instead, the modulation of reox and acid/base properties of the metal centre upon changes in protonation state in the second coordination sphere may be the actual origin of improved catalytic performance.¹⁴⁻¹⁷

Regarding catalysts for H₂ formation, the N₂P₂ ligands introduced by DuBois and co-workers,⁵⁻⁷,¹² and the azaditiolate (adt) bridging ligand in diiron complexes of the general formula Fe₂(adt)(CO)₄X₄ (X = e.g. CO, R₃P) modelled after the FeFe-Hase active site,⁹,²²⁻²⁴ are prominent examples of ligand motifs with basic sites in the second coordination sphere. The amine function of the adt ligand is believed to assist enzymatic H₂ formation and activation by shuttling of protons to or from the distal iron centre of the active site where the crucial terminal hydride intermediate is formed (Scheme 1a).²⁵,²⁶ A similar role of the adt ligand in catalysis by the synthetic models considered here seems less likely given the preferred bridging coordination of the hydride (Scheme 1b).²³ The superior catalytic performance of Fe₂(adt)(CO)₆, 1, over its propyldithiolate (pdt)
analogue 2 in electrochemical H₂ formation has however been attributed to proton shuttling, involving specifically terminal protonation of the metal centre by tautomerization of a adt-NH⁺ precursor upon one-electron reduction of the catalyst.²⁷ We were therefore intrigued whether the proposed role of the adt ligand as proton shuttle and the formation of terminal hydride intermediates in these model complexes could be inferred from direct spectroscopic observation. For this purpose we have combined laser flash induced reduction as well as rapid chemical reduction of the catalyst with both UV-vis and IR detection to elucidate structure and reactivity of reduced and reduced-protonated intermediates derived from 1 (ref. 28) and Fe₂(adt)(CO)₆(PMe₃)₂ (4).²⁹ This approach enabled us to exclude any proton shuttling role of the adt bridging ligand in catalytic H₂ formation while the kinetics of ligand and metal protonation steps suggest an alternative rationale for the improved performance.

Results and discussion

Generally, complex 1 catalyses H₂ formation via an initial ET-PT sequence with weaker acids, or a PT-ET sequence when stronger acids are employed that initially protonate the adt-N with a pKₐ of 8 in acetonitrile²⁸ (Scheme 2, potentials from ref. 27).

Reduction of 1

To investigate the ET-PT route, the one-electron reduced catalyst 1⁻ was obtained by rapid electron transfer from laser flash generated [Ru(dmbs)]⁺ according to eqn (1) and (2) (kET₁ = 3 × 10⁹ M⁻¹ s⁻¹, eqn (2)). The electronic absorption bands of 1⁻ (570, 700 nm, see ESI†) and the carbonyl region of its IR spectrum (1915, 2005 cm⁻¹, Fig. 1a) resemble closely the spectra reported for the singly reduced pdt analogue 2⁻.₃₁,₃₂ The 700 nm band and the uniform shift of the three ν_C=O bands by 80–90 cm⁻¹ towards lower wavenumbers are distinct markers of the intact (μ²,κ³-adt)Fe₂(CO)₆ core in the reduced complex 1⁻, in contrast to the dissociation of a sulphur–iron bond in the reduced bdt analogue (bdt = benzene-1,2-dithiolate) 3⁻.₃₃ In the absence of proton sources, decay of 1⁻ leads to complete recovery of 1 by diffusion controlled charge recombination with the electron donor radical TTF⁺ (krec₁ = 6 × 10¹⁰ M⁻¹ s⁻¹, eqn (3)).

[Ru(dmbs)]⁺ + TTF⁻ ⇌ [Ru(dmbs)]⁺ + TTF⁺ (1)

[Ru(dmbs)]⁺ + I⁻ ⇌ [Ru(dmbs)]²⁺ + I⁻ (2)

1⁻ + TTF⁺ ⇌ [Ru(dmbs)]⁺ + TTF⁺ (3)

Metal vs. ligand protonation of 1⁻

Protonation of the laser flash generated 1⁻ (eqn (4)) with weak acids (Cl₃CCOOH, pKₐ = 10.6 or ClCH₂COOH, pKₐ = 15.3),³⁴

Fig. 1 (a) Transient IR spectrum 1 μs after excitation (○, ∙) and normalized IR spectrum of the starting state (—) for the reduction of 1 (1.3 mM) by flash-quench generated [Ru(dmbs)]⁺ in acetonitrile. (b) Kinetic traces (….) and fits (——) monitoring the pseudo-first order formation of 1⁻ (1915 cm⁻¹) and bleach of 1 (2035 cm⁻¹) followed by second order charge recombination with TTF⁺.
which avoids protonation of the neutral parent complex, shifts all three \( \nu_{C=O} \) bands by about 20 cm\(^{-1} \) to higher wavenumbers (2025, 1965 and 1935 cm\(^{-1} \)), as shown in Fig. 2. The magnitude of the shift is similar to what is observed upon protonation of the neutral parent complex with stronger acids (cf. Fig. 3a) and characteristic for protonation of the adt-N.

\[
\text{I}^- + \text{AH} \rightarrow \text{I}H + \text{A}^- \tag{4}
\]

Metal protonation of \( \text{I}^- \) is on the other hand expected to shift the carbonyl bands by about 80 cm\(^{-1} \) to higher wavenumbers. Shifts of this magnitude have been established for the bridging hydrides formed by protonation of \( \text{Fe}_2(\text{I},\text{I}) \) complexes with electron donor ligands like 4' (ref. 35) as well as transiently generated \( \text{Fe}_2(\text{I},0) \) hexacarbonyl complexes 2' and 3'.\(^{32,33} \) Mutually cancelling shifts of the carbonyl bands upon one-electron reduction and metal protonation were further evidenced with help of complex 4 (\textit{vide infra}). The transient IR spectrum of the protonation product hence proves that the adt ligand remains the preferred site of protonation, at least kinetically, also upon one-electron reduction of \( \text{I} \). This conclusion is corroborated by the UV-vis spectrum of 1H (ESI†), which is similar to its precursor 1', while metal protonation of 2' was previously shown to cause complete bleach of its visible absorption bands that arise from transitions involving predominantly metal-based orbitals. The observed pseudo-first order rate constants for formation of 1H follow a linear dependence on acid concentration before becoming limited by

the preceding electron transfer step (ESI†). The straightforward concentration dependence excludes any complications of the protonation kinetics due to dimerization of the acetic acids that was previously found to result in quadratic concentration dependence corresponding to an increase in effective acid strength.\(^{34} \) Interestingly, protonation of 1', which occurs unambiguously at the adt-N, proceeds remarkably slowly with bimolecular rate constants \( k_{\text{PT}} \) of \( 4 \times 10^8 \text{M}^{-1} \text{s}^{-1} \) and \( 2 \times 10^7 \text{M}^{-1} \text{s}^{-1} \) for \( \text{Cl}_3\text{CCOOH} \) and \( \text{CH}_2\text{H}_4\text{CCOOH} \), respectively.

Notably, the protonation rate constant is almost two orders of magnitude below the diffusion controlled limit even for the strongly exergonic reaction with \( \text{Cl}_3\text{CCOOH} (\Delta pK_a = 3) \). This is quite untypical for protonation of an amine base and indicative of significant reorganization associated with ligand protonation.

**Spectral comparison against the one-electron reduced hydride 4Hy**

The above assignment of the protonation reaction to ligand protonation builds to a large extent on the general notion of nearly perfectly cancelling shifts in carbonyl frequencies caused by one-electron reduction and formation of a bridging hydride. This expectation could previously only be supported by the vanishing transient absorption observed upon protonation of complexes 2' and 3'.\(^{32,33} \) Here, these assignments could be

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Fig. 2 (a) Transient IR spectrum 1 \( \mu \)s after excitation (\( \bigcirc \)) and normalized IR spectrum of the starting state (\( \bigcirc \)) for formation of 1H by protonation of 1' with \( \text{Cl}_3\text{CCOOH} (13 \text{ mM}) \) following reduction of 1 (1.3 mM) by flash-quench generated \( [\text{Ru(dmb)}_3]^{2+} \) in acetonitrile. (b) Kinetic traces (\( \cdots \)) and fits (\( \cdots \)) monitoring pseudo-first order formation of 1H (1965 cm\(^{-1} \)) and bleach of 1 (2035 cm\(^{-1} \)) followed by second order charge recombination with TTF*.

Fig. 3 (a) FTIR spectra of 4 (1.5 mM) (\( \bigcirc \)) in acetonitrile. 4H\(^+\) (\( \square \)) obtained by protonation with toluenesulfonic acid (3 mM), and 4Hy\(^+\) (\( \triangle \)) obtained by protonation with toluenesulfonic acid (3 mM) in presence of tetrabutyl-ammonium chloride (1.5 mM) after correction of the raw spectrum (\( \bigtriangleup \)) for minor contribution from 4H\(^+\). (b) Transient IR spectrum upon reduction of 4Hy\(^+\) to 4Hy by electron transfer from flash-quench generated \([\text{Ru(dmb)}_3]^{2+} (\bigcirc, 1 \mu \text{s after excitation})\) and normalized IR spectra of 4Hy\(^+\) (\( \square \)) and 4 (\( \bigcirc \)).
eventually verified by the positive observation of a one-electron reduced hydride 4Hy by laser flash induced reduction of hydride complex 4Hy'. The latter can be obtained by the Cl⁻ catalysed protonation of 4 that yields the thermodynamically stable hydride instead of the kinetically preferred ligand protonated 4H⁺ in analogy to the previously reported behaviour of 4'. Upon one electron reduction of 4Hy' the carbonyl bands of the product 4Hy coincide almost perfectly with those of the parent complex 4 (Fig. 3) verifying the close spectral resemblance of the diiron carbonyl complexes and their bridging hydride derivatives in the Fe₂(1,0) oxidation state.

Probing possible tautomerization of 1H

According to the above arguments, also formation of a possibly more stable hydride 1Hy by tautomerization (5) should lead to complete decay of all transient absorption given the predictable spectral similarity between 1 and 1Hy.

$$1H \xrightarrow{k_{taut}} 1Hy$$

While the expected spectral changes are the same as for charge recombination between 1H and TTF⁺ (6), tautomerization should nevertheless accelerate the decay of 1H if it proceeds with a rate that is at least comparable to charge recombination (6).

$$1H + TTF⁺ \xrightarrow{k_{rec}} 1 + TTF + H⁺$$

In flash photolysis experiments, 1H decays with second order kinetics on a time scale of about 10 μs in transient IR and about 100 μs with the lower transient concentrations in UV-vis experiments. These time scales are typical for diffusion controlled recombination ($k_{rec,2}$ = $1.5 \times 10^{10}$ M⁻¹ s⁻¹) and put an upper limit on the order of 10⁵ s⁻¹ on the rate constant of a possibly competing tautomerization reaction.

As a turnover frequency (TOF) of 10⁵ to 10⁷ s⁻¹ has been proposed from electrocatalytic experiments (log($k_{cat}$/s⁻¹) = 3.9 and 2.6 for 6 mM Cl₂C₃COOH and CH₂C₃COOH, respectively), we wanted to probe the reactivity of 1H on a longer time scale than 100 μs. Thus, 1H⁺ was chemically reduced by cobaltocene in IR stopped flow experiments (rapid-scan FTIR). 1H⁺ was obtained by addition of two equivalents of 2,5-dichlorobenzenesulfonic acid (Cl₂BSA, pKₐ = 6.7) to warrant quantitative protonation of 1. Reduction to 1H occurred on a shorter time scale than the time resolution of the experiment. Importantly, 1H was observed for several seconds (Fig. 4). Its decay with a lifetime of 1.3 s yields a product spectrum that can be assigned to the parent complex 1 or the tautomerization product 1Hy given the expected similarity of their spectra. The former could be generated by catalytic turnover that leads to depletion of acid and therefore regenerates unreduced, unprotonated catalyst. Whether the slow decay of 1H is predominantly due to turnover or tautomerization cannot be discriminated but the observed lifetime demonstrates in either case that the rate constant for tautomerization, if this occurs at all, has to be smaller than 1 s⁻¹. We note that tautomerization would be an intramolecular process that is independent on external conditions, such as identity and concentration of acid and the method of reduction (chemical or electrochemical).

Reduction and subsequent protonation of 1H⁺

Alternative formation of 1H via the PT-ET route was studied by laser-flash induced reduction of 1H⁺ (7).

$$[Ru(dmb)₃]^{−} + 1H⁺ \xrightarrow{k_{ex}} [Ru(dmb)₃]^{2+} + 1H$$

1H⁺ was obtained by protonation of 1 with Cl₂BSA that shifts the IR peaks by about 20 cm⁻¹ to higher wavenumbers (2090, 2051 and 2015 cm⁻¹). The transient IR absorption spectrum after reduction of 1H⁺ (Fig. 5) shows the bleaching of all three ν_{C-C} peaks of 1H⁺ together with the same product absorption peaks previously attributed to 1H (2025, 1965, 1935 cm⁻¹). While the ET-PT and PT-ET sequences lead expectedly to the same product 1H, its decay is much faster when generated in presence of the stronger acid required for the initial ligand protonation. The accelerated decay of 1H (Fig. 5) without appearance of new transient signals can be assigned to its metal protonation yielding 1HHy⁺ (8).

$$1H + AH \xrightarrow{k_{cat}} 1HHy⁺ + A⁻$$

This assignment is based on the expected IR absorption of 1HHy⁺ that is likely to cancel the bleach of absorption from the starting material 1H⁺ as previously found for hydrides 2Hy and 3Hy. Protonation is also manifested in a slower charge recombination of 1HHy⁺ with TTF⁺ (ESI†). Kinetic analysis of both 1H and TTF⁺ signals results in a protonation rate constant of $k_{PT2} = 6 \times 10⁷$ M⁻¹ s⁻¹ (see ESI†). The analysis eliminates the possible effect of accelerated recombination with accumulated TTF⁺ in the presence of acid. With a rate constant of $6 \times 10⁷$ M⁻¹ s⁻¹ formation of the hydride 1HHy⁺ by metal protonation of 1H with Cl₂BSA (pKₐ = 6.7) occurs about as fast as metal protonation of the pdt analogue 2⁻ ($7 \times 10⁷$ M⁻¹ s⁻¹) by TsOH.
Implications for (electro)catalytic H₂ formation

From previous electrochemical experiments comparing 1 and 2 it has been concluded that catalytic H₂ formation from the same acids proceeds with similar second order catalytic rate constants despite the lower overpotential arising from the easier reduction of adt-protonated catalyst.²⁷ Undiminished reactivity towards protonation despite the lowered driving force would imply that 1 breaks the usual scaling relation between overpotential and rate,²⁸ an effect that has been attributed to the proton shuttling role of the adt²⁷ ligand. By putting an upper limit of 1 s⁻¹ on the first order rate constant for intramolecular tautomerization, our results exclude however that a mechanism involving hydride formation by proton shuttling at the Fe₂(0,0) redox level could under any conditions result in the reported pseudo-first order catalytic rate constant on the order of 10⁸ to 10⁹ s⁻¹, as previously proposed.²⁷ It is important to note that the reported kinetics refer to the first catalytic wave (−1.4 V) that has been attributed to mechanisms where the hydride is formed prior to the second electron transfer, i.e. a CECE sequence with stronger acids or an ECCE sequence with weak acids. Therefore, possible formation of the hydride via tautomerization at the Fe₂(0,0) redox level (1H⁻ → 1Hy⁺) cannot explain the superior catalytic performance of the adt complex.

An alternative advantage of the adt ligand might instead arise from the lowered barrier for hydride formation due to the structural changes at the iron core induced by the preceding ligand protonation. This notion would be in line with the relative sluggishness of the latter reaction (see above) and is corroborated by structural data that has been previously obtained from EXAFS spectra²⁹ of ligand- and metal-protonated states of the related Fe₂(s,i) complex 4. Specifically, protonation of the adt ligand in 4 results in an elongated Fe-Fe bond distance in 4H⁺ that is close to the distance found in the bridging hydride 4H⁻H²⁺. Formation of the hydride from the ligand protonated complex requires hence much less reorganization of this coordinate than formation of hydride 4H⁻H²⁺ from 4.

Generally, a proton shuttling ligand could accelerate not only the formation of the hydride intermediate but also its subsequent coupling with a proton. In case of adt complex 1, with its bridging hydride intermediate, neither of the two steps is, however, likely to involve the adt-N, for steric reasons. In contrast to the terminal hydride of the enzyme active site, the hydride is bridging in the model complex (cf. Scheme 1), as evidenced by the conserved symmetry shown by the IR spectra, and the absence of a bridging CO IR band. The superior catalytic performance of 1 over its pdt analogue 2 is instead attributed to the observed acceleration of hydride formation with bulk acid as described above.

Conclusions

In summary, we could demonstrate by real-time spectroscopic observation that the initial protonation of hydrogenase model complex 1, also in its one-electron reduced Fe₂(1,0) state 1⁻, occurs on the adt ligand rather than the Fe₃ core. The same ligand protonated intermediate 1H was alternatively generated by one-electron reduction of 1H⁻ and no evidence for tautomerization of 1H to a potentially more stable hydride 1Hy was observed in either case. 1H was stable on the time scale of 1 s, which shows that intramolecular proton shuffling cannot be responsible for the reported catalytic TOPs of 10⁸ to 10⁹ s⁻¹.²⁷ For catalytic H₂ generation from weak acids this implies that no hydride intermediate is formed on the Fe₂(0,0) level; instead, hydride formation may occur after further reduction, by direct protonation in the Fe₂(0,0) state. Only with strong acids a doubly protonated intermediate 1H²⁺ is formed in the Fe₂(0,0) state, by direct metal protonation of 1H, not by ligand

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Fig. 5 (a) Transient IR spectrum 0.3 µs after excitation (○, ) and normalized IR spectrum of the starting state ( — ) for formation of 1H by reduction of 1H⁺ (obtained by protonation of 1 with 2.6 mM Cl₂BSA) with flash-quench generated [Ru(dmb)]⁺ in acetonitrile. (b) Kinetic traces ( ••••••••) and fits ( — — — — ) monitoring pseudo-first order formation of 1H (2051 cm⁻¹) and bleach of 1H⁺ (2051 cm⁻¹) followed by pseudo-first order decay of 1H by protonation with Cl₂BSA and charge recombination with TTF⁺.

(pKᵢ 8–8.7).³⁴ Both complexes are in the same Fe₂(1,0) oxidation state but the pKᵢ of the hydride should drop by about 5 units according to a shift of +330 mV in reduction potential for the 1H⁻/1H couple compared to 1/1⁻ or 2/2⁻. Protonation of 1H, even by the stronger Cl₂BSA, is therefore thermodynamically less favourable than protonation of 2⁻ by TsOH, with ΔpKᵢ lowered by about 3 for the former reaction. Applying the free energy relationship (Brønsted coefficient of 0.4) established for protonation of 2⁻,³¹ the protonation of 1H to 1H⁺H²⁺ is hence one to two orders of magnitude faster than expected for protonation of its pdt analogue 2⁻ with the same driving force. Intrinsically faster protonation of the adt-protonated 1H would prevent this step from becoming rate limiting when reducing weaker acids. Below, we discuss the possible origin of this favourable effect.

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protonation of a hydride precursor. Generally, it can therefore be concluded that, in contrast to the enzymatic reaction, the adt ligand of model complex 1 has no proton relay function that enables rapid formation of the bridging hydride intermediates on the Fe$_2$(I,0) level. The protonation kinetics reveal however an exceptionally high barrier for protonation of the adt ligand that presumably due to rearrangements of the metal core upon ligand protonation. In this way the basic site in the second coordination sphere might allow the catalyst to overcome the usual trade-off between overpotential and rate without invoking any proton-shuttling role.

**Conflicts of interest**

There are no conflicts to declare.

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**Notes and references**